

& some interesting
 results were obtained
 from the study of
 the influence of the
temperature on the
rate of reaction
between the
two substances
used.

Conclusions

It was found that
 the rate of reaction
 increases with
 increasing temperature.
 This is due to the fact
 that at higher temperatures
 a larger number of
 molecules possess sufficient
 energy to overcome the
 activation energy barrier.
 The results obtained
 are in agreement with
 the Arrhenius equation.
 The activation energy
 for this reaction was
 found to be approximately
 15 kJ/mol.